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the one best answer

for each question. 1.

In the reaction

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$\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow$

$\text{Ca}(\text{NO}_3)_2 + \text{CO}_2 +$

H_2O , how many

milliliters of 0.1015 M

HNO_3 are required to

react with 0.9857

grams of CaCO_3 ? (a)

0.1940

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With Answers

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excess oxygen.
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Analysis exists as the hydrated salt $BaX_{2.2}H_2O$, where X is the halogen. The barium content of the salt can be determined by Questions And Answers For Gravimetric Analysis Question. The ...

Questions And Answers For

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If the gravimetric procedure cannot be used, use the volume displacement procedure for all of the packages in the sample. 100 multiple type question answer for boe

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the one best answer

for each question. 1.

In the reaction

$\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow$

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$\text{Ca}(\text{NO}_3)_2 + \text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of 0.1015 M HNO_3 are required to react with 0.9857 grams of CaCO_3 ? (a) 0.1940 mL (b) 97.04 mL (c) 194.1 mL 2.

Chemistry 213

Assignment No. 2

Name: Summer 2010

The following quiz contains 5 multiple

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choice questions. If you wish to take a longer quiz, please select 'Review

Questions' from the navigation bar. ...

Which of the following is not a property required of the substance chosen for use as a precipitate in a gravimetric analysis? A hydrated magnesium salt,

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MgSO₄ · xH₂O, was heated until all the ...

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Choice have a molar mass of
180 gmol⁻¹.

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questions which

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with multiple answer options. Primarily, multiple choice questions can have single select or multi select answer options. These are the most fundamental questions of a survey or questionnaire where the

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**Questions And
Answers For
Gravimetric
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CHEM 201 Self Quiz
– 2 (Gravimetric

analysis/Volumetric
analysis) Answer key

1. Calcium soluble
products for SrF₂
given of
saturated solution is
 8.6×10^{-4} M. SrF₂!Sr

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$2+ + 2F^-$ K sp = $[Sr^{2+}][F^-]^2 = 8.6 \times 10^{-4}$
 $[F^-] = 2[Sr^{2+}] = 2 \times 8.6 \times 10^{-4} = 17.2 \times 10^{-4}$
 $[Sr^{2+}] = 8.6 \times 10^{-4} / 17.2 \times 10^{-4} = 1.5 \times 10^{-4}$

2. Calcium fluoride, CaF_2 , is a sparingly ...

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