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Lab 2

Significant  
Lab 2

Figures In Data

# Significant Figures In Data

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Figures - A Fast

Review! Rounding

*2 - Significant*

*Figures* How to

*Page 4/48*

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round to  
significant  
figures (s.f.)

---

Unit Conversion

\u0026

Significant

Figures: Crash

Course Chemistry

#2How To Round A

Number Off To 2

Significant

Figures

(rounding to

2sf) Significant

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*Figures Made*

*Easy! Chemistry*

*Lesson:*

*Significant*

*Digits \u0026*

*Measurements*

~~Significant~~

~~Figures~~

~~Addition~~

~~Subtraction~~

~~Multiplication~~

~~Division \u0026~~

~~Scientific~~

~~Notation Sig~~

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~~Figs The Maths  
Prof: Rounding  
(to 3~~

~~significant  
figures)~~

~~Significant  
figures |~~

~~Decimals | Pre-  
Algebra | Khan  
Academy~~

**Scientific  
Notation and  
Significant  
Figures (1.7)**

*Page 7/48*

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**Significant**

**Digits** *Metric*

*Conversion*

*Trick!! Part 1*

~~The Maths Prof:~~

~~Rounding~~

~~Decimals~~ *Sig Fig*

*Rules!*

*(Significant*

*Figures Rules*

*and Examples)*

*Rounding Using*

*Significant*

*Figures -*



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~~Decimals Adding  
and Subtracting  
Significant  
Figures!~~

---

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Figures:

Multiplication  
and Division!

Precision,

Accuracy and

Uncertainty in  
measurement in  
chemistry

---

Math Antics -

*Page 9/48*

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~~Rounding Examples  
of significant  
figures, when do  
zeros count?~~

**Significant  
Figures**

*Measurement and  
significant  
figures*

*Measurement and  
Significant*

*Figures How to  
Round Numbers to  
Significant*

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Figures #20

~~Precision,~~

~~Accuracy,~~

~~Measurement, and~~

~~Significant~~

~~Figures Sig Fig~~

~~rules~~

~~(Significant~~

~~Figures)~~

Measurement 2-

Significant

Figures

Significant

Figures and Zero

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~~(1.3) Rounding  
to significant  
figures Lab 2~~

~~Significant  
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Lab # 2:

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Calculation:

2.10 cm x 2.20

cm x 1.90cm=

8.78 More digits

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(6 or 7):

8.77800 Correct  
number of

digits: 8.78 (3  
digits) 9.

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Why? The number of digits, i.e. significant figures, reported for a numerical quantity conveys to the reader the precision of the instrument used to make the measurement. In this course when recording data

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in the  
laboratory you  
will have to  
record your  
measurements in  
a way that

~~Lab # 2:~~

~~Significant~~

~~Figures in Data~~

~~— Rust Science~~

Calculation:

2.10 cm x 2.20

cm x 1.90cm=

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## Lab 2

8.78 More digits

(6 or 7) :

8.77800 Correct

number of

digits: 8.78 (3

digits) 9. Post

Lab questions a)

Why can't we

write numbers

with as many

significant

figures as we

want?



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Lab

~~#2 — Significant  
figures —~~

~~DavidHonChemPer8~~

Unit II:

Measurement and

Significant

Figures LAB

REPORT I.

Purpose: The

purpose of this

laboratory

experiment is to

obtain

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measurements involving mass and volume and to perform calculations using correct significant figures. Four experiments will be performed using lab equipment from the eScience Lab Kit and

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dimensional  
analysis will be  
applied to  
calculations  
that involve  
converting  
units.

~~Lab 2.docx~~

~~Unit II~~

~~Measurement and~~

~~Significant~~

~~Figures ...~~

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research, as  
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difficulty as  
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straightforward

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here. As this

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significant

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~~Significant~~

~~Figures In Data~~

For example, if

your average is

3.025622 and

your standard

deviation is



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0.01845, then this is the correct number of significant figures for the average: 3.03, because the first digit of the standard deviation is in the hundredths place, so the last significant digit of the

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average is in  
the hundredths  
place.

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Figures Lab +  
Middlebury  
College Chem 103  
lab~~

You simply  
include all the  
significant  
figures in the  
leading number.

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For example, the number 450 has two significant figures and would be written in scientific notation as  $4.5 \times 10^2$ , whereas 450.0 has four significant figures and would be written as  $4.500 \times 10^2$ . In scientific

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notation, all significant figures are listed explicitly.

~~2.4: Significant Figures in Calculations — Chemistry ...~~

1.423 x 4.2 =  
6.0 since 1.423  
has 4  
significant

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figures and 4.2  
only has two  
significant  
figures, the  
final answer  
must also have 2  
significant  
figures.  $234.67$   
 $- 43.5 = 191.2$   
since 43.5 has  
one decimal  
place and  $234.67$   
has two decimal  
places, the

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final answer

must have just  
one decimal

place.

~~Significant~~

~~Figures +~~

~~Introduction to~~

~~Chemistry~~

Example: 356

rounded to 2

significant

digits is 3 6 0.

This calculator

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rounds down if the next digit is less than 5 and rounds up when the next digit is greater than or equal to 5. In the table below 305.459 is rounded from 0 to 6 significant figures. For comparison the same number is

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rounded from 0  
to 6 decimal  
places.

~~Rounding~~

~~Significant~~

~~Figures~~

~~Calculator~~

How to use the  
sig fig  
calculator. Our  
significant  
figures  
calculator works



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in two modes -  
it performs  
arithmetic  
operations on  
multiple numbers  
(for example,  
 $4.18 / 2.33$ ) or  
simply rounds a  
number to your  
desired number  
of sig figs.  
Following the  
rules noted  
above, we can

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calculate significant figures by hand or by using the significant figures counter.

~~Significant  
Figures  
Calculator — Sig  
Fig~~

Following are the significant figures rules that govern the

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determination of significant figures: Those digits which are non-zero are significant. For example, in 6575 cm there are four significant figures and in 0.543 there are three significant figures. If any

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zero precedes the non-zero digit then it is not significant.

~~What are the~~

~~Rules for~~

~~Significant~~

~~Figures—~~

~~Precision ...~~

example: Round to 2 significant figures:  $2.35 \times 10^2$  (Answer:

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$2.4 \times 10^2$ )

example: Round  
to 2 significant

figures:  $2.45 \times$   
 $10^2$  (Answer:

$2.4 \times 10^2$  ) Of  
course, if we

round to 2  
significant

figures:  $2.451 \times$   
 $10^2$  , the

answer is

definitely  $2.5 \times$   
 $10^2$  since  $2.451$

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$2.5 \times 10^{-2}$  is closer  
to  $2.5 \times 10^{-2}$   
than  $2.4 \times 10^{-2}$

.

~~Math Skills~~

~~Scientific~~

~~Notation~~

For example, the  
number 450 has  
two significant  
figures and  
would be written  
in scientific

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notation as  $4.5 \times 10^2$ , whereas 450.0 has four significant figures and would be written as  $4.500 \times 10^2$ . In scientific notation, all significant figures are listed explicitly.

Example 2.4.1

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~~2.4: Significant~~

~~Figures in~~

~~Calculations~~

~~Chemistry ...~~

Title:

Measurement Lab

2 Significant

Figures

Calculations 1

Measurement Lab

2 Significant

Figures

Calculations 2.



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Objectives ; Use proper technique to make accurate and precise measurements.

Apply the rules for significant figures to calculations.

Informal

Assessment ;

Monitoring

student

interactions and

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questions as

Figures In Data

~~PPT~~

~~Measurement Lab~~

~~2 Significant~~

~~Figures~~

~~Calculations ...~~

Lab Quiz: 1.

Round off the

following

measurement to

three

significant

figures. 1.296 g

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### Significant Figures in Data

2. How many significant figures are there in the following measurement?

2020 g 3. Round off the following measurement to three significant figures. 5.658 grams

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~~Virtual Lab~~

~~Precision and~~

~~Significant~~

~~Figures – Mr ...~~

Rules about significant figures may seem arbitrary from a theoretical standpoint, but in the laboratory you will see that

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they allow you to determine the precision of your measurements and calculations.

When your measurement has a limited number of digits, your subsequent calculations will also have a limited number

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of digits.

Figures In Data

~~Significant~~

~~Figures +~~

~~Middlebury~~

~~College Chem 103~~

~~lab~~

In

multiplication

and division the

number of

significant

figures is

simply

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determined by the value of lowest digits.

This means that if you multiplied or divided three numbers: 2.1, 4.005 and 4.5654, the value 2.1 which has the fewest number of digits would mandate

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that the answer  
be given only to  
two significant  
figures.

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