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Calculation : Basic Chemical

Principles Material Balance

Problem Approach CHEMICAL

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Mnitrogen+ 0.21 Moxygen. =

$0.79(28) + 0.21(32) = 28.84$ The

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mass fractions of oxygen and nitrogen in air are then $m_{\text{oxygen}} = (0.21)(32)/28.84 = 0.233$, or 23.3% and $m_{\text{nitrogen}} =$

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4. Write additional data required to solve the problem and the chemical equations if the process involves

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5. Select a suitable basis of calculations. 6.

List by symbols each of the unknown values of the stream flows and compositions 7. Make a number of independent material balances equations equal to unknown

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Basic Principles and Calculations in
Chemical Engineering

Lec 1: Definition, History, Role of
Chemical Engineer; Lec 2: Basic
Features of Chemical Process; Lec
3: Unit systems and dimensions;
Process Variables and Rate. Lec 4:

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Calculations and Properties of Notes

Material in System; Lec 5:

Pressure and Temperature of Flow

Process; Lec 6: Rate of Process;

Fundamentals of Material Balance

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view the study of the field of

chemical engineering as a tree with material and energy balances being the trunk and the subjects of thermodynamics, fluid flow, heat transfer, mass transfer, reactor kinetics, process control, and process design being the branches

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off the trunk. From this perspective, it is

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$z + zP. z. z = \hat{g} (2.2.8)$ Setting
the limit on the left hand side as
 $z \rightarrow 0$ gives the definition of the

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derivative of P with respect to z:

$$dP/dz = -\rho g \quad (2.2.9)$$

In other words, the change in pressure, P, with height, z, is negative – pressure reduces with height (or increases with depth – as any diver knows).

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Part 1. Fundamental Concepts in
Chemical Engineering Calculations

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Lecture 2. Regression and

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Fundamental Process Variables

Lecture 4. The Chemical Equation and Stoichiometry Part 2. Gasses and Phase Equilibrium

Lecture 5. Ideal Gas Calculations

Lecture 6. Real Gas Relationships

Lecture 7. Vapor-Liquid Equilibria Part 3.

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