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protons in the nucleus do
not change during normal
chemical reactions. Only the
outer electrons move.

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7. Write a balanced chemical equation for each chemical reaction: a) Solid zinc reacts with aqueous hydrochloric acid (HCl) to form zinc chloride and hydrogen gas. b) Gaseous

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methane (CH_4) reacts with oxygen gas to form carbon dioxide gas and liquid water.

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of the following options
correctly interpret the
symbol ${}^{224}_{86}\text{Rn}$? 1. Rn is

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the symbol of the element 2.
The number 86 is the atomic
number of the isotope

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chemical bond. cation.
anion. atoms tend to gain,
lose, or share electrons in
order to acqui... the force
that holds two atoms
together; may form by the
attra... positively charged
ion. negatively charged ion.

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Chemical Reactions 1

Although chemists recognize
that matter can be converted

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into energy and energy into matter, this matter?energy conversion is small enough to be disregarded. 7.1

Energy253 objective7

objective5 objective7

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The p Block Elements.

Question 1. Elements in the groups 13 to 18 in the Periodic table constitute the 'p' block elements.

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Chemical Equilibrium In a
chemical reaction chemical
equilibrium is defined as
the state at which there is

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no further change in concentration of reactants and products. For example,

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Choice Questions. Question
1. The equilibrium
expression, $K_c = \frac{[CO_2]}{[CO]^2}$
represents the reaction.
Question 2. Hydrogen

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molecule (H_2) can be dissociated into hydrogen atoms (H). Which one of the following changes will not increase the number of atoms present at equilibrium?

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help students understand the
concept very well.

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Questions. Question 1.
Discuss the general characteristics of group 15 elements with reference to their electronic configuration, oxidation state, atomic size, ionisation enthalpy and

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electro negativity. Answer:
General characteristic of
group - 15 elements : (A)
Electronic configurations :

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Chemistry Chapter 7 - The p-
Block Elements The p-Block
Elements consists of groups
13 to 18 of the periodic
table. Chapter 7 chemistry
class 12 is a continuation
of class 11 syllabus. While

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the course of grade 11 deals with the first two groups of p-Block, 13 and 14; class 12 syllabus deals with groups 16 to 18.

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