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Using Gibbs Free Energy The
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Entropy and Second Law of
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Thermochemical Equations
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energy, spontaneity and

equilibrium Thermodynamics:

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Using the second law of
thermodynamics, explain why
heat flows from a hot body to a
cold body but not from a cold
body to a hot body. One test of
the spontaneity of a reaction is
whether the entropy of the

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universe increases: $\Delta S_{\text{univ}} > 0$.

Using an entropic argument, show that the following reaction is spontaneous at 25°C:

19.E: Chemical Thermodynamics (Exercises) - Chemistry ...

chapter. Lecture Outline 19.1

Spontaneous Processes^{1,2,3,4} □

Chemical thermodynamics is concerned with energy

relationships in chemical

reactions. □ We consider

enthalpy. □ We also consider

entropy in the reaction. □ Recall

the first law of thermodynamics:

energy is conserved. $\Delta E = q + w$

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Example 9.2 The element

mercury, Hg, is a silvery liquid at

room temperature. The normal

freezing point of mercury is

-38.9°C , and its molar enthalpy of

fusion is $\Delta H_{\text{fusion}} = 2.29 \text{ kJ/mol}$.

What is the entropy change of the

...

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Terms in this set (28) Enthalpy, Entropy. What makes a process, or more specifically, a chemical reaction, spontaneous? 1.

Tendency for a system to go to a lower state of energy (change in _____ will be a ...

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Concepts: Terms in this set (38)
1) The first law of
thermodynamics can be given as
_____. A) $\Delta E = q + w$ B) $\Delta = -$ C)
for any spontaneous process, the
entropy of the universe increases

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Phase Change Consider the

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melting of ice □ At 1 atm and 0 °C ice and water are in equilibrium with each other. □ If we melt one mole of ice at 0 °C/1 atm to form 1 mole of water still at 0 °C □ We need to give the water heat that would be equal to ΔH_{fusion} . □ If the process is done extremely slowly the process

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Thermodynamics □ Enthalpy □ A
thermodynamic quantity that
equal to the internal energy of a
system plus the product of its
volume and pressure exerted on
it by its surroundings; "Enthalpy
is the amount of energy in a
system capable of doing
mechanical work" Using the
symbol H for the enthalpy: ...

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Chapter 19. Chemical Thermodynamics . 19.2 Entropy and the Second Law of Thermodynamics . Entropy Change □ Entropy, S , is a thermodynamic term that reflects the disorder, or randomness, of the system. □ The more disordered, or random, the system is, the larger the value of S .

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equal to the internal energy of a
system plus the product of its
volume and pressure exerted on
it by its surroundings; "Enthalpy
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system capable of doing
mechanical work" Using the
symbol H for the enthalpy:

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Chemical thermodynamics
(thermochemistry) is the study of
energy flow during a chemical
reaction. ... There are 3 "laws of
thermodynamics". In Chapter 6,
you learned the first law
(conservation of energy: there is
a constant amount of energy in

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