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Name: Kaelyn Hoffman The Combined

Gas Law • Expresses the relationship

between the pressure, volume and

temperature of a fixed amount of gas.

• $PV/T = K$ or $P_1V_1/T_1 = P_2V_2/T_2$ Ex:

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A sample of gas has a volume of 201 L when its temperature is 293 and its pressure is 24 mm Hg. What volume will the gas occupy at STP?

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Q. Deep-sea divers use a breathing mixture of helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen if the total pressure of the mixture is 760 mm Hg, the pressure exerted by the helium is 609 mm Hg, and the pressure exerted by the carbon dioxide is 1 mm Hg?

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Think about these steps to answer the question: Identify the system - P.E.). goes from liquid PE change, In this This is an type of change, the system (the water) heat from the surroundings. Identify the surroundings - This reaction is When Act E is high, the reaction is (slow, fast). Sketch a diagram of these reactions: thermic. faster, exothermic

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9-08 Series circuit: In a series circuit:

Current - Equation: Voltage - Equation:

Resistance - Equation: Ex Problem: R

$T = \frac{V}{R}$ $V = IR$ $I = \frac{V}{R}$ $V = IR$

$V_1 = IR_1$ $V_2 = IR_2$ $V_3 = IR_3$

Problem Set #3 | $I_1 = I_2 = I_3 = I$

Note Taking Guide: Program 903 - Part 2 Name

Season 1 Episode 903 | 29m 59s

More About the Behavior of Gases:

The Combined Gas Law, Graham's

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